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## INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

<b>(51) International Patent Classification <sup>6</sup> :</b> <b>C03C 10/16, 3/247</b>	<b>A1</b>	<b>(11) International Publication Number:</b> <b>WO 99/05071</b> <b>(43) International Publication Date:</b> 4 February 1999 (04.02.99)
<b>(21) International Application Number:</b> PCT/US98/12262 <b>(22) International Filing Date:</b> 11 June 1998 (11.06.98)  <b>(30) Priority Data:</b> 60/053,588                      24 July 1997 (24.07.97)                      US  <b>(71) Applicant (for all designated States except US):</b> CORNING INCORPORATED [US/US]; 1 Riverfront Plaza, Corning, NY 14831 (US).  <b>(72) Inventor; and</b> <b>(75) Inventor/Applicant (for US only):</b> TICK, Paul, A. [US/US]; 10615 Lennox Drive, Corning, NY 14830 (US).  <b>(74) Agents:</b> HERZFELD, Alexander, R. et al.; Corning Incorporated, Patent Dept., SP FR 02-12, Corning, NY 14831 (US).		<b>(81) Designated States:</b> AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GE, GH, HU, ID, IL, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ZW, European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE).  <b>Published</b> <i>With international search report.</i>
<b>(54) Title:</b> TRANSPARENT LANTHANUM FLUORIDE GLASS-CERAMICS  <b>(57) Abstract</b>  The present invention relates to a transparent glass-ceramic material which includes a glass matrix and a crystalline phase of lanthanum fluoride crystals in the glass matrix, where the transparent glass-ceramic material contains no silica. The present invention further relates to a method of producing the transparent glass-ceramic material.		

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## TRANSPARENT LANTHANUM FLUORIDE GLASS-CERAMICS

### FIELD OF THE INVENTION

5           The present invention relates to a transparent glass-ceramic material and a method of making the material.

### BACKGROUND OF THE INVENTION

10           Recently, transparent materials capable of efficient frequency upconversion, most being various rare-earth ion-doped fluoride glasses and crystals, have received great attention due to the possibilities of utilizing these materials to achieve blue or green solid  
15           state lasers. While no significant difference in upconversion efficiency is observed between fluoride glasses and single crystals, single mode optical fiber doped with a low level of rare-earth ions can be drawn from fluoride glasses, bringing about highly efficient  
20           blue or green upconversion fiber lasers. Unfortunately, heavy metal fluoride glasses suffer certain undesirable attributes which have restricted their applications. Most notably, heavy metal fluoride glasses exhibit poor resistance to devitrification. U.S. Patent No. 4,674,835  
25           to Mimura et al. discusses the crystallization problems of heavy metal fluoride glasses, one example of which is called ZBLAN, and the light scattering problems resulting therefrom.

30           The great susceptibility of heavy metal fluoride glasses to devitrification also generates problems in forming large preforms. Crystallization at

the interface between the core and cladding, during the production of the preform, causes problems in the most commonly used methods for preparing an optical fiber. That is, heavy metal fluoride glasses are quite prone to inhomogeneous nucleation, the consequence of which being crystallization at the core and cladding interfaces, particularly during the drawing of the optical fiber. The resulting fibers are subject to serious scattering losses due to crystals in the fibers.

Devitrification of the heavy metal fluoride glasses is aggravated when ions necessary to impart differences in indices of refraction to the core and cladding are added to the glass composition. Additional doping, for example, with rare-earth metal ions, also tends to reduce the stability of the glass. As a consequence of those problems, research has focused on finding additives which will reduce the tendency of the glass to devitrify and to increase the chemical stability thereof. In addition, the preparation of fluoride glasses requires the glass forming components to be reheated to their softening temperatures, which generally are about 75° C above the glass transition temperatures. In addition, fluoride glasses cannot be melted in air but require a water-free, inert gas environment.

Most oxide glasses (such as silica dioxide) are much more chemically and mechanically stable, are easier to prepare, and are more easily fabricated into rods, optical fibers, or planar waveguides than fluoride glasses. Unfortunately, due to their larger phonon energy, silica glasses are very inefficient for infrared upconversion. It has also been shown that addition of oxides into some fluoride glasses improve their stability, but this is not preferred, since even a small addition of oxides will significantly quench the upconversion luminescence.

Early in 1975, Auzel et al., J. Electrochem. Soc., 122:101 (1975) reported an interesting class of infrared ("IR") upconversion materials which were prepared from classical glass forming oxides ( $\text{SiO}_2$ ,  $\text{GeO}_2$ ,  $\text{P}_2\text{O}_5$ , etc. with  $\text{PbF}_2$  and rare-earth oxides), and showed an efficiency nearly twice as high as  $\text{LaF}_3\text{:Yb:Er}$  phosphor. Since these kinds of materials were comprised of a mixture of glassy and crystalline phases, and the embedded crystals were very large in size (around 10  $\mu\text{m}$ ), they were not transparent.

In Wang et al., "New Transparent Vitroceramics Codoped With  $\text{Er}^{3+}$  and  $\text{Yb}^{3+}$  For Efficient Frequency Upconversion," Appl. Phys. Lett., 63(24):3268-70 (1993), transparent oxyfluoride vitroceramics (also called glass-ceramics) containing oxides of large phonon energy, like  $\text{SiO}_2$  and  $\text{AlO}_{1.5}$ , but showing IR to visible upconversion more efficient than fluoride glass was described. The composition of Wang contained, expressed in terms of mole percent,

$\text{SiO}_2$	30	$\text{CdF}_2$	20
$\text{AlO}_{1.5}$	15	$\text{YbF}_3$	10
$\text{PbF}_2$	24	$\text{ErF}_3$	1

The glass produced from that composition was heat treated at 470°C to develop nanocrystallites which the authors stated did not reduce the transparency of the body. The authors posited that the  $\text{Yb}^{3+}$  and  $\text{Er}^{3+}$  ions were preferentially segregated from the precursor glass and dissolved into the nanocrystals upon heat treatment. The size of the nanocrystallites was estimated by the authors to be about 20 nm; that size being so small that light scattering loss was minimal. The authors reported the upconversion efficiency of their products to be about 2

to 10 times as high as that measured on the precursor glass and other fluoride-containing glasses. However, the crystals which are formed in the Wang glass have a cubic lattice structure and this limits the concentration of some of the trivalent rare-earth elements which may be incorporated into the crystal phase. Another problem with these materials is that they require cadmium in the formulation. Cadmium is a carcinogen and, thus, its use is restricted. Hence this type of glass would not be desirable for any large scale manufacturing operation.

The present invention is directed toward overcoming these above-noted deficiencies.

#### SUMMARY OF THE INVENTION

The present invention relates to a transparent glass-ceramic material which includes a glass matrix and a crystalline phase of lanthanum fluoride crystals in the glass matrix, where the transparent glass-ceramic material does not contain silica.

Another aspect of the present invention relates to a method of making a transparent glass-ceramic material which includes providing a glass matrix and treating the glass matrix under conditions effective to produce the transparent glass-ceramic material which includes a glass matrix and a crystalline phase of lanthanum fluoride crystals in the glass matrix, where the transparent glass-ceramic material does not contain silica.

Transparent glasses containing crystals are highly desirable in applications where there is a requirement for the glass to be easily melted or formed and additionally to contain a crystal, where the crystal itself may be difficult or expensive to synthesize. Such transparent glasses, containing lanthanum fluoride crystals, are especially desirable where the crystal itself provides highly desirable features, such as

optical activity. In addition, the glass-ceramics of the present invention are easily produced and can be melted at low temperature relative to many silicate glasses, in air. Further, lanthanum fluoride has the particular advantage of being able to accommodate large concentration of any rare-earth element into its crystal structure and is not a carcinogen.

#### BRIEF DESCRIPTION OF THE DRAWINGS

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Figure 1 is a graph comparing the emission spectra of a glass-ceramic of the present invention with a fluoride glass at 1550 nm.

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Figure 2 is an X-ray diffraction pattern of a precursor glass of the present invention.

Figure 3 is an X-ray diffraction pattern of a glass-ceramic of the present invention.

20

Figure 4 is a Differential Thermal Analysis ("DTA") curve showing the position of the glass transition temperature ( $T_g$ ) and the crystallization temperature required to produce a glass-ceramic of the present invention.

25

Figure 5 is a diagram showing the optimum region for forming a glass-ceramic of the present invention.

#### DETAILED DESCRIPTION OF THE INVENTION

30

The present invention relates to a transparent glass-ceramic material which includes a glass matrix and a crystalline phase of lanthanum fluoride crystals in the glass matrix, where the transparent glass-ceramic material does not contain silica.

35

Another aspect of the present invention relates to a method of making a transparent glass-ceramic material, which includes providing a glass matrix and treating the glass matrix under conditions effective to

produce the transparent glass-ceramic material, which includes a glass matrix and a crystalline phase of lanthanum fluoride crystals in the transparent glass-ceramic material, where the glass-ceramic material does not contain silica.

Preferably, the transparent glass-ceramic materials include lanthanum fluoride crystals having uniform size distribution and interpartical separations of the order of 10-40 nanometers. Further, it is preferable for the size of each crystal of the crystalline phase to be in the range of about 5-15 nanometers. Further, the glass-ceramic material of the present invention preferably contains from about 5% to about 30% by volume crystals. Thus, an "ultra-transparent" glass-ceramic is produced, which is useful for doping with active rare-earth elements. "Ultra-transparent" signifies that the glass-ceramic material of the present invention, when doped with active rare-earth elements, has optical scattering properties which are indistinguishable from a glass, in particular, having minimal light scattering loss.

It is desirable for the transparent glass-ceramic composition to include 45 to 55 wt.%  $\text{LaF}_3$ , 20-25 wt.%  $\text{B}_2\text{O}_3$ , 10-20 wt.%  $\text{Al}_2\text{O}_3$ , and 10-15 wt.%  $\text{BaF}_2$ . Optionally, the composition may also include 0-15 wt.%  $\text{PbF}_2$ , 0-10 wt.%  $\text{Y}_2\text{O}_3$ , and 0-10 wt.%  $\text{Ta}_2\text{O}_5$ .

In addition, the transparent glass composition may include fluorides or oxides of one or more rare-earth elements other than lanthanum. Preferably, the composition includes from 0 - 5 wt. % fluorides or oxides of these additional rare-earth elements. The additional rare-earth elements are Y, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, and Lu.

The rare-earth elements may be contained entirely in the crystalline phase. Alternatively, the rare-earth elements are contained both in the glass matrix and in the crystalline phase.



The transparent glass-ceramic material is produced by providing a glass matrix and treating the glass matrix under conditions effective to produce the transparent glass-ceramic material, which includes a glass matrix and a crystalline phase of lanthanum fluoride crystals in the glass matrix, wherein the glass-ceramic material does not contain silica.

Optionally, the glass matrix is provided as a formed glass shape. Alternatively, the glass matrix is provided by producing the glass matrix. The glass matrix is produced by any traditional glass making procedure. For example, glass components which include alkoxides and are formulated to produce a glass matrix having the desired composition are used in a sol gel approach, such as disclosed in U.S. Patent No. 5,494,863 to Mathur, which is hereby incorporated by reference, to produce the glass matrix. In addition, for example, the glass matrix may be produced using chemical vapor deposition techniques.

A particularly preferred way of providing the glass matrix is by producing the glass matrix through traditional melting and forming techniques. Preferably, the glass matrix is produced by first melting together batch components calculated to produce a glass matrix having the desired composition to give a glass melt. Importantly, no silica is utilized. Preferably, the glass components are melted at temperatures from about 1000 C to about 1200 C for from about .25 to about 2 hours. Next, the glass melt is formed into a glass shape. Suitable forming procedures include rolling, pressing, casting, or fiber drawing. The glass shape is then preferably a patty, rod, sheet, or fiber.

Subsequently, the glass matrix provided is treated under conditions effective to produce a transparent glass-ceramic material having a glass matrix and a

crystalline phase of lanthanum fluoride crystals in the glass matrix.

Preferably, the treating step is achieved by heating the glass matrix to ceram the glass matrix. The heating step is designed to promote phase separation and crystallization of the lanthanum fluoride phase, producing a glass-ceramic material which includes a crystalline phase of lanthanum fluoride in a glass matrix. Preferably, the glass matrix is heated in a heating step at a temperature from about 600°C to about 700°C. The exact heat treatment temperature is determined by the crystallization behavior of the lanthanum fluoride phase, which can be determined by differential thermal analysis. The glass matrix is then cooled to room temperature. Perferably, the article is rapidly cooled to the anneal temperatures (about 560°C), then cooling proceeds at a rate sufficient to remove residual stresses.

The transparent glass-ceramic material is especially useful where a glass matrix, which is easily melted or easily formed, is desired or where the glass matrix contains a crystalline phase which may be difficult or expensive to produce.

Further, a transparent glass-ceramic material which includes lanthanum fluoride crystals in the glass matrix, is especially useful where the crystals act as selective hosts for other rare-earth elements. A dramatic change is observed in the emission spectra of  $\text{ErF}_3$  doped lanthanum fluoride glasses after ceramming, indicating significant partitioning of the rare-earth element into the crystal phase.

It is not necessary to have complete partitioning of the rare-earth element into the crystalline phase. For certain applications, such as gain-flattened amplifiers, for example, a combination of  $\text{Er}^{3+}$  in the glass matrix and in the crystalline phase may

be optimal. This permits tailoring of the site sizes and geometries of the lanthanum fluoride crystals to optimize the local environments of any dopant lanthanide ions.

#### EXAMPLES

##### 5 Example 1

Various precursor glass materials were produced as follows. Glass forming batch materials calculated to produce the desired compositions were mixed, and the mixture was melted in covered platinum crucibles for 1-  
10 2 hours at 1200°C in air to produce a glass melt. The glass melt then was cast to produce a precursor glass.

Examples of representative precursor glass compositions obtained are listed in Table 1 below. Precursor glasses having the appearance of clear glass  
15 are desirable for ceramming to produce the transparent glass-ceramic compositions of the present invention.

Thus, compositions 1-3, 7, 11, 18, 21, 27-30, 34, 35, 41, and 42 are desirable precursor glasses. Figure 5 illustrates an optimum region for forming  
20 desirable precursor glasses.

Table 1

ID	La <sub>2</sub> O <sub>3</sub>	LaF <sub>3</sub>	B <sub>2</sub> O <sub>3</sub>	PbF <sub>2</sub>	PbO	Al <sub>2</sub> O <sub>3</sub>	Y <sub>2</sub> O <sub>3</sub>	ZnF <sub>2</sub>	BPO <sub>4</sub>	BaF <sub>2</sub>	SiO <sub>2</sub>	Ta <sub>2</sub> O <sub>5</sub>	Appearance
1	50		25									25	clear glass
2	37.5	12.5	25									25	clear glass
3	12.5	37.5	25									25	clear glass
4		50	25									25	crystals/glass
5		50	25		25								crystals/glass
6		40	25		25						10		devitrified
7		40	25		25	10							clear glass
8		50	20	10							10		phase separated
9		50	20		25	5							devitrified
10		50	25		15	10							hazy glass
11		50	25		10	15							clear glass
12		50	25			25							didn't melt
13		50	25		5	15		5					crystals glass
14		50	25		5	20		5					didn't melt
15		50	25			15		10					partly melted
16		50	20		15	15							surface crystals/ glass
17		50	15		25	10							devitrified
18		50	25	5	5	15							clear glass
19		50	25	10		15							devitrified
20		45	25	10		15							5 ErF <sub>3</sub> clear glass
21		50	25	10		15							clear glass

ID	La <sub>2</sub> O <sub>3</sub>	LaF <sub>3</sub>	B <sub>2</sub> O <sub>3</sub>	PbF <sub>2</sub>	PbO	Al <sub>2</sub> O <sub>3</sub>	Y <sub>2</sub> O <sub>3</sub>	ZnF <sub>2</sub>	BPO <sub>4</sub>	BaF <sub>2</sub>	SiO <sub>2</sub>	Ta <sub>2</sub> O <sub>5</sub>	Appearance
22		50		10		15			25				didn't melt
23		50	25	10		7.5	7.5						surface devit.
24		50	20	15		7.5	7.5						devitrified
25		50	25			7.5	7.5	10					devitrified
26		50	20	10		15			5				devitrified
27		50	25	10		10				5			clear glass
28		50	25	10		5				10			clear glass
29		50	25	5		10				10			clear glass
30		50	25			10				15			clear glass
31		50	20			10				15	5		surface devit.
32		50	20			10				20			surface devit.
33		50	22.5			15				10			2.5 AgF, devit.
34		50	25			10				15			clear glass
35		50	25			10				15			clear glass
36		50	22.5			17.5				15			clear glass
37		50	22.5			12.5	5			10			few crystals
38		50	20			12.5	7.5			10			clear glass
39		50	22.5			10	5			10			surface devit.
40		50	22.5			5	5			10			2.5 AlF <sub>3</sub> , s. crystals
41		50	20			10	5			12.5			5 AlF <sub>3</sub> , s. crystals
42		50	15			10	5			10		5	clear glass
												10	clear glass

Example 2

A transparent glass-ceramic material with lanthanum fluoride crystals having a composition as shown in Table 2 below was prepared by first preparing the precursor glass as described in Example 1. The precursor glass then was cerammed at 680°C to produce a transparent glass-ceramic material.

Table 2

LaF <sub>3</sub>	47.5 wt. %
B <sub>2</sub> O <sub>3</sub>	25 wt. %
PbF <sub>2</sub>	10 wt. %
Al <sub>2</sub> O <sub>3</sub>	15 wt. %
Er <sub>2</sub> O <sub>3</sub>	2.5 wt. %

A comparison of the emission spectra of this sample (line 1) versus the emission spectra of a pure fluoride glass (line 2) at 1550 nm is shown in Figure 1. Importantly, the transparent glass-ceramic has a long, very flat plateau, extending to 1560 nm, indicating that this material is useful as an amplifier host, because of the uniform gain in this area.

Example 3

A transparent glass-ceramic having lanthanum fluoride crystals having a composition shown in Table 3 below was prepared as described in Example 2.

Table 3

LaF <sub>3</sub>	50 wt. %
B <sub>2</sub> O <sub>3</sub>	20 wt. %
Al <sub>2</sub> O <sub>3</sub>	10 wt. %
PbO	20 wt. %

As shown in Figure 2, after melting and before ceramming, X-ray diffraction indicates that the precursor material was a glass, with no crystalline phase present. As shown in Figure 3, after ceramming, a glass-ceramic was present, having a crystalline LaF<sub>3</sub> phase.

Example 4

A transparent glass-ceramic having lanthanum fluoride crystals having a composition shown in Table 4 below was prepared.

Table 4

LaF <sub>3</sub>	50 wt. %
B <sub>2</sub> O <sub>3</sub>	25 wt. %
Al <sub>2</sub> O <sub>3</sub>	15 wt. %
PbF <sub>2</sub>	10 wt. %

As shown in Figure 4, a Differential Thermal Curve indicates the glass transition temperature (T<sub>g</sub>) of about 550°C for the composition and the ceramming temperature (T<sub>c</sub>) of approximately 700°C. Specifically, the onset T<sub>g</sub> was 567°C, with a midpoint of 575°C. At a temperature of approximately 890°C, additional undesirable crystallization occurs, indicating that lower ceramming

temperatures are necessary to produce the glass-ceramic material of the present invention.

5 The optimum region for forming a four component system of the inventive glass-ceramic is shown in Figure 5. In this model, the glass ceramic is comprised of 50 wt.%  $\text{LaF}_3$ , and 50 wt.%  $(\text{B}_2\text{O}_3 + \text{BaF}_2 + \text{Al}_2\text{O}_3)$ . The region indicated in Figure 5 represents the optimum region for components  $\text{B}_2\text{O}_3$ ,  $\text{BaF}_2$  and  $\text{Al}_2\text{O}_3$ .

10 Although the invention has been described in detail for the purpose of illustration, it is understood that such detail is solely for that purpose, and variations can be made therein by those skilled in the art without departing from the spirit and scope of the invention which is defined by the following claims.



WHAT IS CLAIMED:

- 1        A transparent glass-ceramic material  
comprising:
- 5                a glass matrix and  
                 a crystalline phase of lanthanum fluoride  
crystals in the glass matrix, wherein the transparent  
glass-ceramic material contains no silica.
- 10              2        The transparent glass-ceramic material  
according to claim 1, wherein the lanthanum fluoride  
crystals have a uniform size distribution.
- 15              3        The transparent glass-ceramic material  
according to claim 2, wherein the lanthanum fluoride  
crystals have a size of no greater than about 15  
nanometers.
- 20              4        The transparent glass-ceramic material  
according to claim 1, wherein the transparent glass-  
ceramic material comprises: 45 to 55 wt. %  $\text{LaF}_3$ , 20 - 25  
wt. %  $\text{B}_2\text{O}_3$ , 10 - 20 wt. %  $\text{Al}_2\text{O}_3$ , and 10 - 15 wt. %  $\text{BaF}_2$ .
- 25              5        The transparent glass-ceramic material  
according to claim 4 further comprising: 0 - 15 wt. %  
 $\text{PbF}_2$ , 0 - 10 wt. %  $\text{Y}_2\text{O}_3$  and 0 - 10 wt. %  $\text{Ta}_2\text{O}_5$ .
- 30              6        The transparent glass-ceramic material  
according to claim 4, further comprising:  
                 up to 5 wt. % fluorides or oxides of one  
or more additional rare-earth elements.
- 35              7        The transparent glass-ceramic material  
according to claim 6, wherein the additional rare-earth  
elements are selected from the group consisting of Y, Ce,  
Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, and Lu.

8       The transparent glass-ceramic material according to claim 7, wherein the additional rare-earth elements are contained entirely in the crystalline phase.

5               9       The transparent glass-ceramic material according to claim 7, wherein the additional rare-earth elements are contained both in the glass matrix and in the crystalline phase.

10              10       A method of making a transparent glass-ceramic material comprising:  
                    providing a glass matrix and  
                    treating the glass matrix under conditions effective to produce the transparent glass-ceramic  
15       material comprising a glass matrix and a crystalline phase of lanthanum fluoride crystals in the glass matrix, wherein the transparent glass-ceramic material contains no silica.

20              11       The method according to claim 10, wherein the transparent glass-ceramic material comprises: 45 to 55 wt.%  $\text{LaF}_3$ , 20-25 wt.%  $\text{BaF}_2$ , 10-20 wt.%  $\text{Al}_2\text{O}_3$ , and 10-15 wt.%  $\text{BaF}_2$ .

25              12       The method according to claim 11, wherein the transparent glass-ceramic material further comprises: 0-15 wt. %  $\text{PbF}_2$ , 0-10 wt.%  $\text{Y}_2\text{O}_3$  and 0-10 wt.%  $\text{Ta}_2\text{O}_5$ .

30              13       The method according to claim 11, wherein the transparent glass-ceramic material further comprises:  
                    up to 5 wt.% fluorides or oxides of one or more additional rare-earth elements.

35              14       The method according to claim 13, wherein the additional rare-earth elements are selected from the

group consisting of Y, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, and Lu.

15        15    The method according to claim 14, wherein  
the additional rare-earth elements are contained entirely  
in the crystalline phase.

10        16    The method according to claim 14, wherein  
the additional rare-earth elements are contained both in  
the crystalline phase and in the glass matrix.

15        17    The method according to claim 10, wherein  
the providing comprises:  
             melting the glass matrix to produce a  
glass melt and  
             forming the glass melt into a glass shape.

20        18    The method according to claim 17, wherein  
the melting is performed at temperatures from about 1000°C  
to about 1200°C for about .25 to 2 hours.

25        19    The method according to claim 10, wherein  
the treating comprises:  
             heating the glass matrix to form the  
transparent glass-ceramic material comprising a  
crystalline phase of lanthanum fluoride crystals in the  
glass matrix.

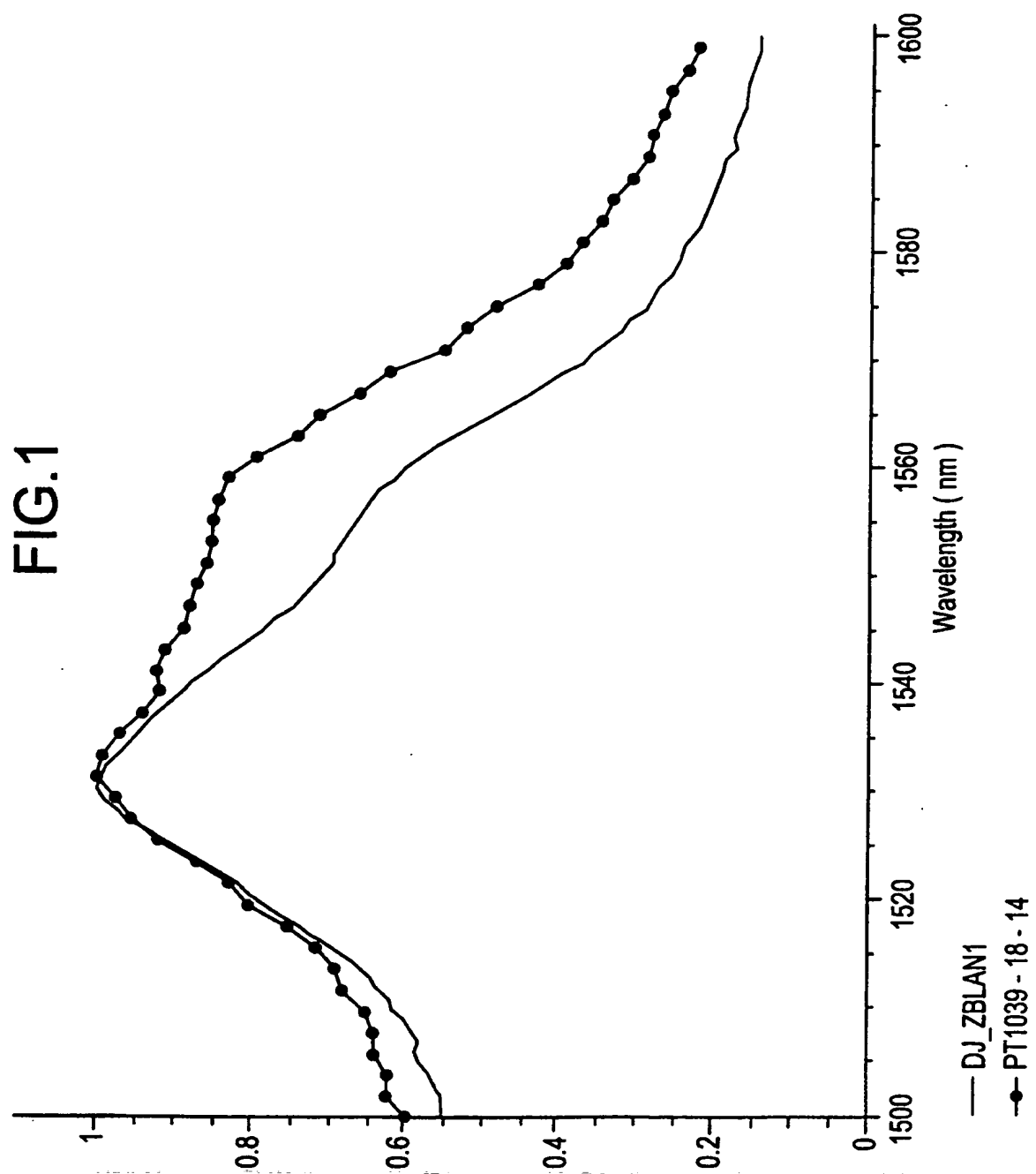
30        20    The method according to claim 19, wherein  
the heating comprises:  
             heating the glass matrix at a temperature  
from about 600°C to about 700°C.

35        21    A product prepared by the method according  
to claim 10.

22 A product prepared by the method according  
to claim 17.

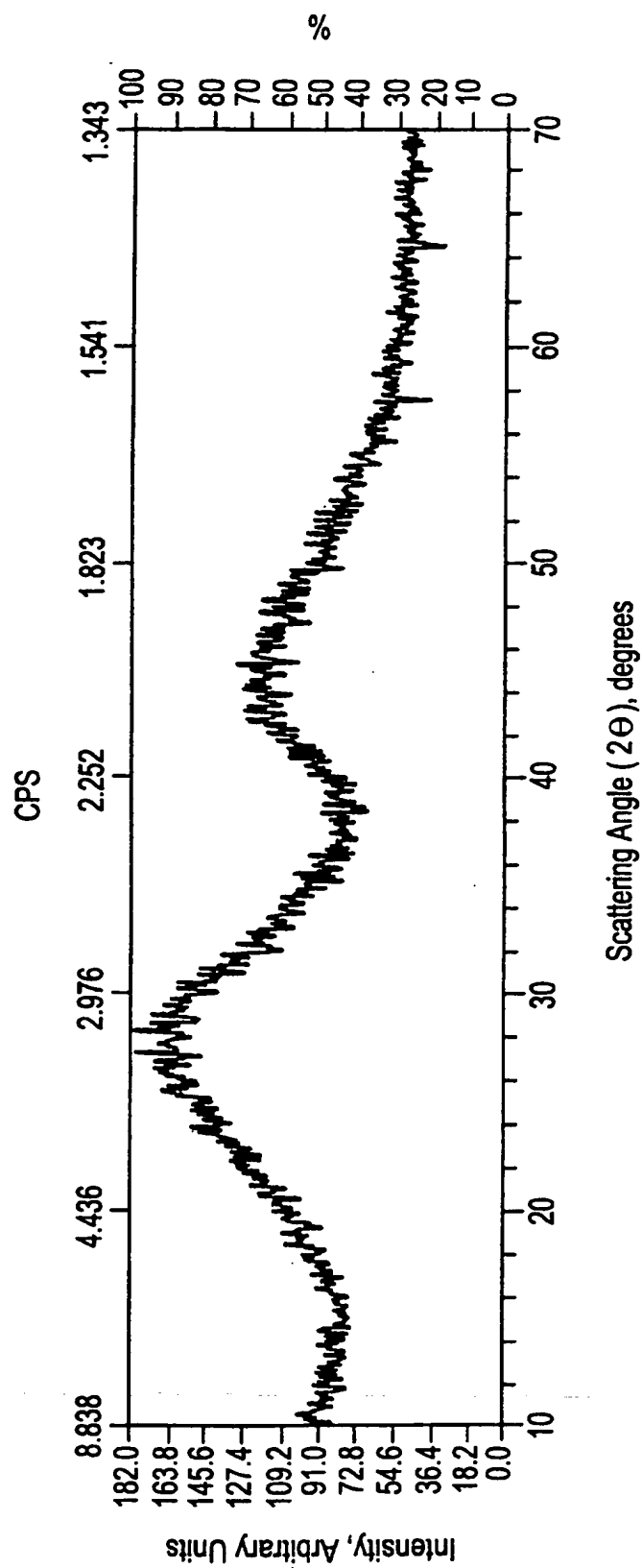
5 23 A product prepared by the method according  
to claim 20.

1 / 5



2 / 5

FIG.2



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FIG.3A

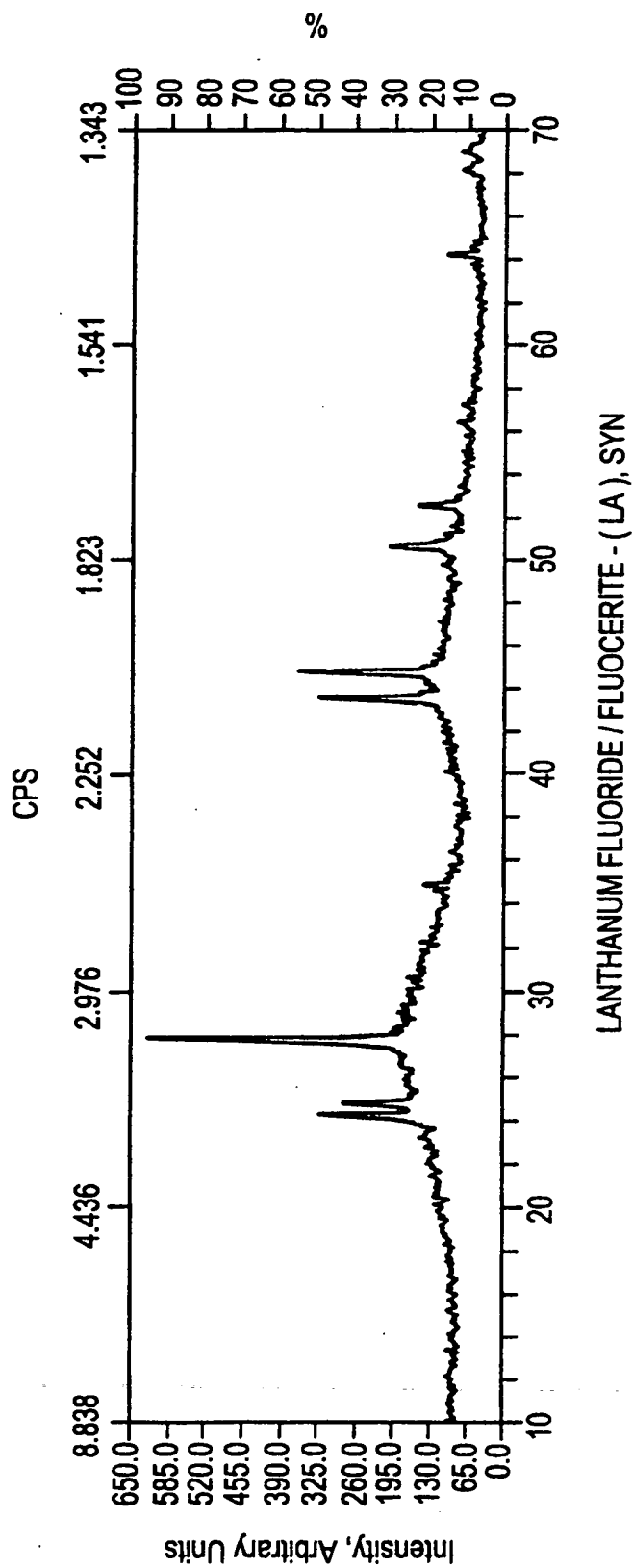
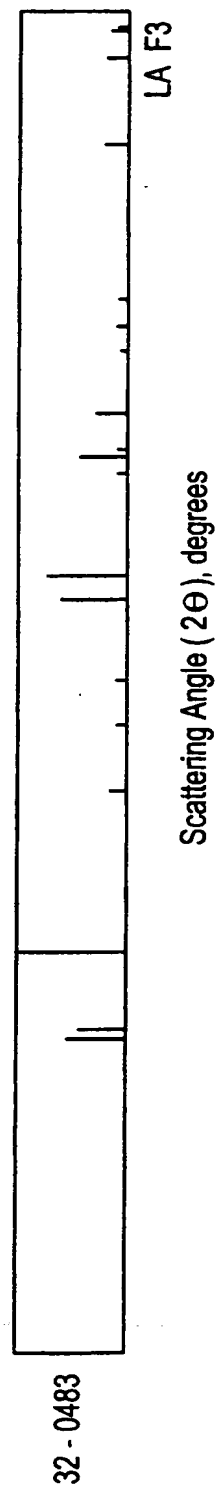
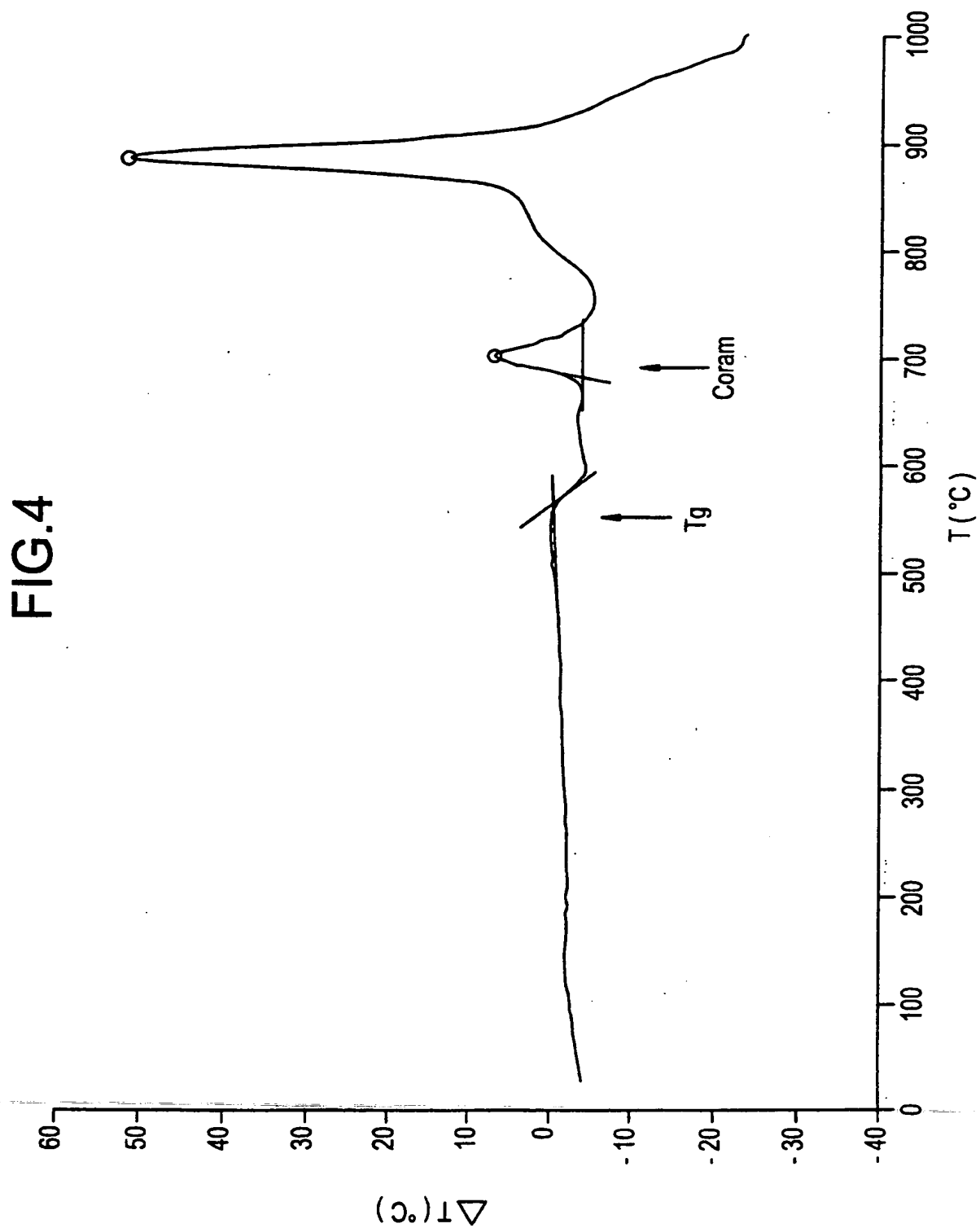


FIG.3B

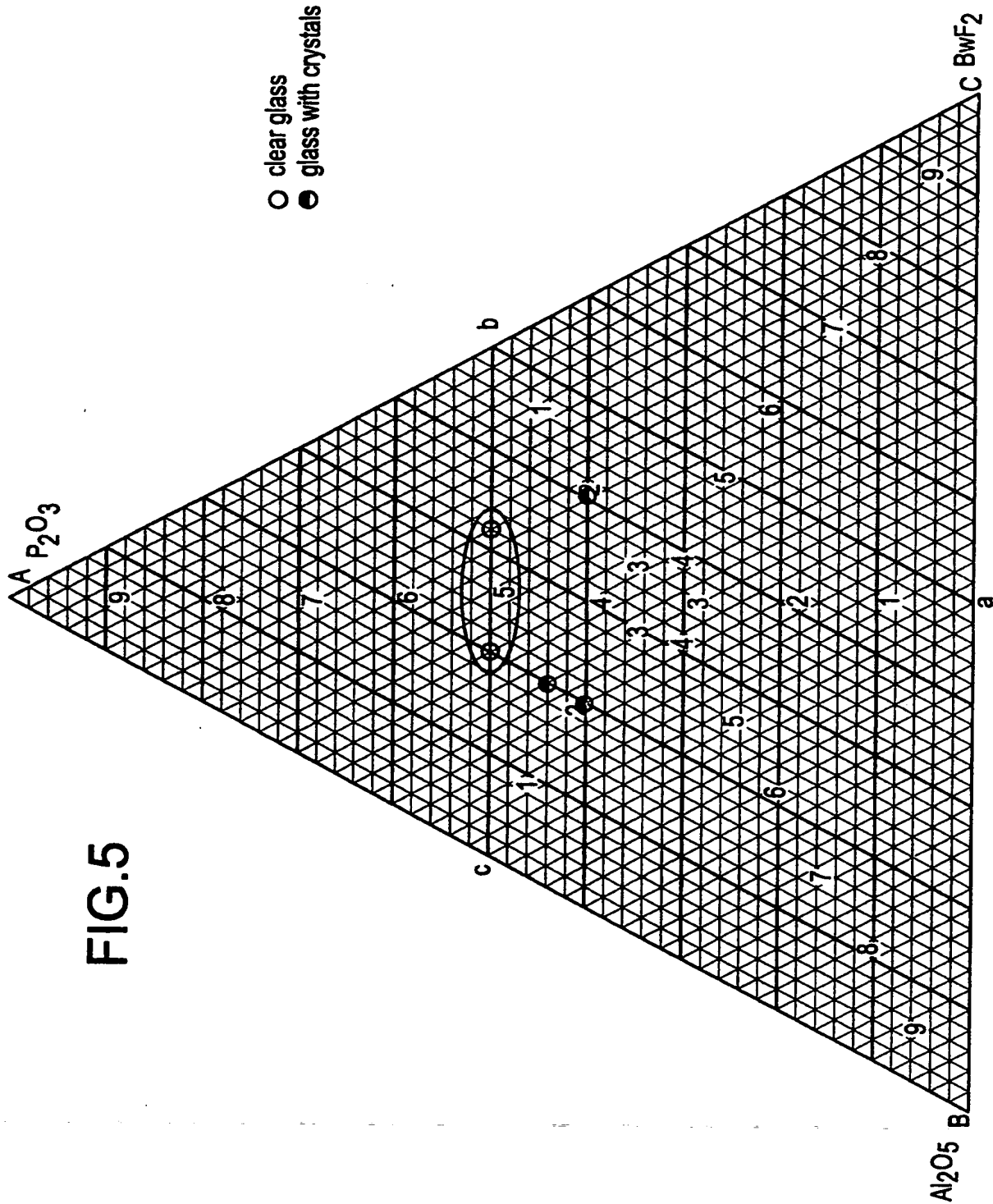


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## INTERNATIONAL SEARCH REPORT

International application No.  
PCT/US98/12262

## A. CLASSIFICATION OF SUBJECT MATTER

IPC(6) : CO3C 10/16, 3/247

US CL : 501/3,44,50,51; 65/33.3

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 501/3,44,50

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	US 4,032,351 A (AUZEL et al) 28 June 1977, col. 3, lines 22-37.	1,210,17,19,21-23
A	US 5,545,595 A (WANG et al) 13 August 1996.	1-23
A	US 5,420,080 A (WANG et al) 30 May 1995.	1-23

☐ Further documents are listed in the continuation of Box C.
 ☐ See patent family annex.

* Special categories of cited documents:	*T* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
*A* document defining the general state of the art which is not considered to be of particular relevance	*X* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
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*O* document referring to an oral disclosure, use, exhibition or other means	
*P* document published prior to the international filing date but later than the priority date claimed	

Date of the actual completion of the international search

03 SEPTEMBER 1998

Date of mailing of the international search report

16 OCT 1998

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